

Middle School



Elements, Compounds, & Mixtures DOODLE NOTES

Name Mrs. Morehouse

Scientist Classify matter into one of three categories:

ELEMENTS

1. Contain only one type of atom.
 2. CANNOT be broken down into simpler substances by either physical or chemical means. N₂ O₂ are diatomic molecules!
 3. Can exist as either atoms or molecules.

COMPOUNDS

1. Contain two or more elements in a specific ratio.
 2. CANNOT be broken down into simpler substances (elements) by physical means.
 3. Can exist as either atoms or molecules.

MIXTURES

1. Contain two or more substances (elements or compounds) that are NOT chemically combined.
 2. Can be separated by physical means.
 3. Retains properties of its components.


You try!
 When magnesium (Mg) metal is burned in the presence of oxygen (O₂), magnesium oxide (MgO) is produced. Is MgO an element, compound, or mixture? Explain.

Salad	Air
Pizza	Muddy water
Kool-Aid	

Filtration
Evaporation

Why use Doodle Notes in *your* science class?

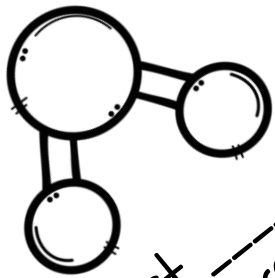
- ✓ Increased focus
- ✓ Better retention of material
- ✓ Stronger learning ability for new content
- ✓ Relaxation, decreased stress levels
- ✓ Activates both sides of the brain simultaneously for cross-lateral learning



Activate your student's creativity by encouraging them to use markers, crayons, or colored pencils to doodle, copy notes, highlight important vocabulary, and add their own notes.

Read more about maximizing communication between both hemispheres of your student's brains on [Math Giraffe's blog](#). It contains interesting research information about the real benefits of including coloring and doodling in science and math.

Name _____



Scientist
into one of three categories:

ELEMENTS

1. Contain only _____ type of atom.
2. _____ be broken down into simpler substances by either _____ or _____ means.
3. Can exist as either _____ or _____



COMPOUNDS

1. Contain _____ of _____ or more elements together in a specific _____
2. _____ be broken down into simpler substances (_____ means but NOT by _____)
3. Has _____ properties than its component _____



MIXTURES

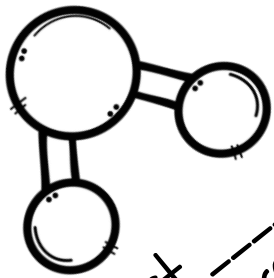
1. Contains _____ or _____ substances (elements together).
2. Can be separated _____
3. Retains many of the same _____ of its components.

You try!



When magnesium (Mg) metal is burned in the presence of oxygen (O₂), magnesium oxide (MgO) is produced. Is MgO an element, compound, or mixture? Explain.

Name _____



Scientist
into one of three categories:
matter

ELEMENTS



COMPOUNDS



MIXTURES

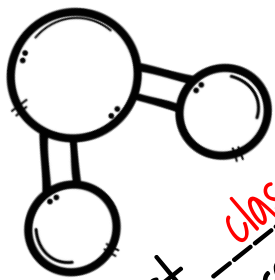


You try!

When magnesium (Mg) metal is burned in the presence of oxygen (O_2), magnesium oxide (MgO) is produced. Is MgO an element, compound, or mixture? Explain.

Name _____

Teacher Key



Scientist classify matter into one of three categories:

ELEMENTS

1. Contain only one type of atom.
2. CANNOT be broken down into simpler substances by either physical or chemical means.
3. Can exist as either atoms or molecules.



Na H ₂ O ₂ F B	

COMPOUNDS

1. Contain atoms of one or more elements bonded together in a specific ratio.
2. CAN be broken down into simpler substances (physical or chemical) means but NOT by physical or chemical elements.
3. Has different properties than its component elements.



H ₂ O NaCl KCl C ₆ H ₁₂ O ₆	

MIXTURES

1. Contains one or more substances (elements or compounds) that are NOT physically bonded together.
2. Can be separated physically.
3. Retains many of the same properties of its components.



salad Kool-Aid pizza air	

When magnesium (Mg) metal is burned in the presence of oxygen (O₂), magnesium oxide (MgO) is produced. Is MgO an element, compound, or mixture? Explain.

It is a compound because it is two elements bonded together.

You try!

SAMPLE

Name Mrs. Morehouse

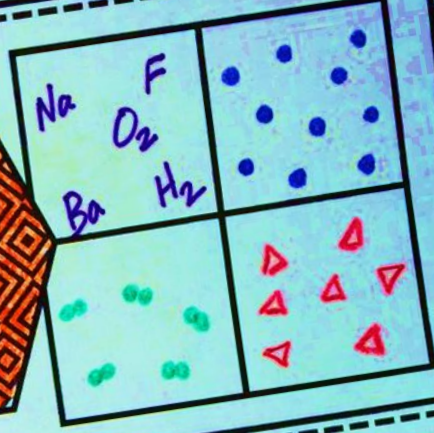


Scientist Classify matter into one of three categories:

ELEMENTS

1. Contain only one type of atom.
2. CANNOT be broken down into simpler substances by either physical or chemical means.
3. Can exist as either atoms or molecules.

N₂ O₂ are diatomic molecules!



COMPOUNDS

1. Contain atoms of one or more elements together in a specific ratio.
2. CAN be broken down into simpler substances (elements) but NOT by physical means.
3. Has different properties than its component elements.

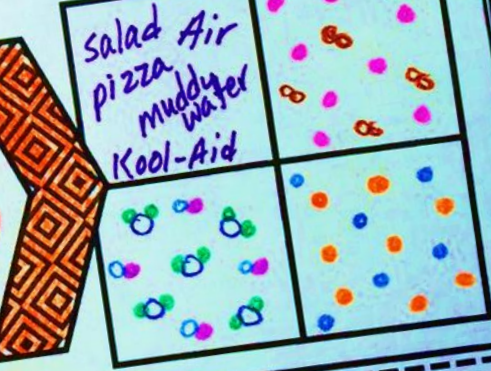
1:2, 2:3, etc. bonded elements



MIXTURES

1. Contains one or more substances (elements or compounds) that are NOT physically bonded together.
2. Can be separated physically.
3. Retains many of the same properties of its components.

You try!



When magnesium (Mg) metal is burned in the presence of oxygen (O₂), magnesium oxide (MgO) is produced. Is MgO an element, compound, or mixture? Explain.
It is a compound because it is two elements that are bonded together.

LOTS OF ROOM FOR NOTES AND EVEN ADDITIONAL NOTES AND DOODLES!

Credits:

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